[www.ck12.org](http://www.ck12.org/) **Voltaic Cells Practice Worksheet**

1. In a voltaic cell, the anode and cathode are placed in two different compartments. ( True/False )
2. In a salt bridge, the electrolytes such as KCl are preferred since their ions have an almost equal transport number. ( True/False )
3. In a voltaic cell, the electrode on which oxidation takes place is called the anode, or negative electrode. ( True/False)
4. Electrons travel from the cathode to the anode (True/False)
5. If a salt bridge is removed from between two half cells, the voltage drops to zero (True/False)
6. A \_\_\_\_\_\_\_\_\_\_\_\_\_\_ cell creates a flow of electrons through a spontaneous reaction
7. Which describes ion migration through the salt bridge or porous membrane?

a. Cations move toward the cathode only.

* 1. Cations and anions both move toward the anode.
	2. Cations move toward the anode and anions move toward the cathode
	3. Cations move toward the cathode and anions move toward the anode
1. Where does oxidation occur in an electrochemical cell?
2. at the cathode, which is the positive electrode
3. at the anode, which is the positive electrode
4. at the anode, which is the negative electrode
5. at the cathode, which is the negative electrode
6. In an electrochemical cell composed of zinc and copper, which correctly describes the reaction taking place in the copper half-cell?
7. the oxidation of copper metal to copper ions
8. the reduction of copper metal to copper ions
9. the reduction of copper ions to copper metal
10. the oxidation of copper ions to copper met

**Answer Keys**

1. True
2. True
3. TRUE
4. False
5. TRUE
6. Galvanic or Voltaic
7. Cations move toward the cathode and anions move toward the anode.
8. at the anode, which is the negative electrode
9. the reduction of copper ions to copper meta